

Name: _____

Band: _____

HW 6.3

Calculate the empirical formula from the following percentages: (You will need to attach a few pages of paper, to neatly show your work.)

1. 63.5% Fe and 36.5% S. = _____

2. 46.6% Fe and 53.4% S. = _____

3. 46.7% N and 53.3% O. = _____

4. 25.9% Fe and 74.1% O. = _____

5. 79.3% W and 20.7% O. = _____

6. 81.8% C and 18.2% H. = _____

7. 31.9% K, 28.9% Cl and 39.2% O. = _____

8. 35.0% N, 4.99% H and 60.0% O. = _____

9. 85.7% C and 14.3% H. = _____

10. 14.6% Mg, 16.9% Si and 68.5% F. = _____

11. 0.83% H, 45.8% Mn and 53.4% O. = _____

Calculate the molecular formula for problems 12, 13, 14, 15:

12. 7.79% C and 92.21% Cl. The molecular mass is 154 grams. = _____

13. 14.46% C and 85.54% Cl. The molecular mass is 166 grams. = _____

14. 10.13% C and 89.87% Cl, the molecular mass is 237 grams. = _____

15. 14.46% C and 85.54% Cl, the molecular mass is 249 grams. = _____