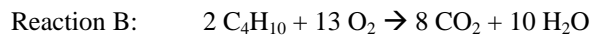
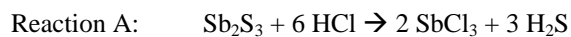


Name: \_\_\_\_\_ Band: \_\_\_\_\_ HW 8.2. "Moles With Grams Stoichiometry."

This problem set on stoichiometry references two different reactions. Each problem references only one of the following balanced reactions. Use the answer spaces, and include the proper units (grams, moles, compound name).



- 1) How many grams of HCl are need to make 4 moles  $\text{SbCl}_3$  ?  
= \_\_\_\_\_ g HCl (enter)\_
- 2) How many grams  $\text{O}_2$  are needed to make 100 g  $\text{H}_2\text{O}$  ?  
= \_\_\_\_\_ g \_\_\_\_\_
- 3) Given 40 g  $\text{Sb}_2\text{S}_3$  , how many moles  $\text{H}_2\text{S}$  are made?  
= \_\_\_\_\_
- 4) Given 92 g  $\text{O}_2$ , how many g  $\text{C}_4\text{H}_{10}$  can be combusted?  
= \_\_\_\_\_
- 5) Given 1 Kg  $\text{C}_4\text{H}_{10}$ , how many g  $\text{CO}_2$  are made?  
= \_\_\_\_\_
- 6) Given 192 g  $\text{Sb}_2\text{S}_3$ , how many g  $\text{H}_2\text{S}$  are made?  
= \_\_\_\_\_
- 7) How many moles  $\text{SbCl}_3$  are made from 84 g HCl?  
= \_\_\_\_\_
- 8) How many grams  $\text{SbCl}_3$  are made from 227 g  $\text{Sb}_2\text{Cl}_3$  ?  
= \_\_\_\_\_
- 9) How many g  $\text{C}_4\text{H}_{10}$  are combusted if 19 moles  $\text{CO}_2$  are made?  
= \_\_\_\_\_
- 10) How many g  $\text{H}_2\text{O}$  are made from 1,100 Kg of  $\text{C}_4\text{H}_{10}$  ?  
= \_\_\_\_\_