

Objective: In this lab, we will observe the five different chemical reactions. Your objective is to determine which reaction you are observing, and to describe the reaction with a predicted chemical equation. Don't worry about balancing these reactions, we'll be using this lab for that process soon.

Compounds and Elements used in this Lab: (You need to name the formulas, first):

Lead II nitrate = _____	Ethanol = <u>C₂H₅OH</u>
Potassium Iodide = _____	Carbon dioxide = _____
Sulfur = (charge, also) _____	Water = _____
Calcium = (charge, also) _____	Oxygen = _____ (diatomic, no charge)
Stannous II Chloride = _____	Hydrogen Peroxide = <u>H₂O₂</u>
Zinc = (charge, also) _____	Clorox is a catalyst, no formula is needed

Reaction 1: In a microplate, mix 4 drops lead II nitrate with 4 drops potassium iodide.

Observation: _____
 _____ --> _____ reaction type: _____

Reaction 2: In a test tube, mix ¼" sulfur powder, and ¼" calcium pieces. Mix, then heat very carefully. You need the assistance of the instructor.

Observation: _____
 _____ --> _____ reaction type: _____

Reaction 3: In a microplate, add several drops of Stannous II Chloride. Add a small piece of Zinc. Prod with a spatula.

Observation: _____
 _____ --> _____ reaction type: _____

Reaction 4: Place a few mL of ethanol into the watch glass. Carefully ignite. Carefully smother the reaction with the lid. Observe what happens on the lid.

Observation: _____
 _____ --> _____ reaction type: _____

How does your observation about what happens on the surface of the lid relate to the reaction above?

Reaction 5: Place 5 drops 3% hydrogen peroxide into a microplate well. Add one drop of clorox bleach, then another etc. Do not spill any bleach on your clothing!

Observation: _____
 _____ --> _____ reaction type: _____