

Name: \_\_\_\_\_ Band: \_\_\_\_\_ L 7.2 Precipitate Reactions

**Objective:** In this lab, we will only observe double replacement reactions. Your objective is to describe the visible products (describe the color and texture of the precipitate) of the following reactions, and to balance these equations. Attach scrap paper to show your work. You will need your polyatomic/common metal charges sheet and a periodic table.

#	Reactants:	Products:	Precipitate description:
1	Mercuric Chloride and sodium carbonate		
2	Nickel (II) Chloride and sodium carbonate		
3	Lead (II) nitrate and potassium iodide		
4	Mercuric chloride and potassium iodide		
5	Barium chloride and sodium sulfate		
6	Cupric chloride and sodium carbonate		
7	Sodium chromate and silver (I) nitrate		
8	Silver (I) nitrate and sodium chloride		
9	Magnesium chloride and sodium hydroxide		
10	Barium nitrate and sodium sulfate		